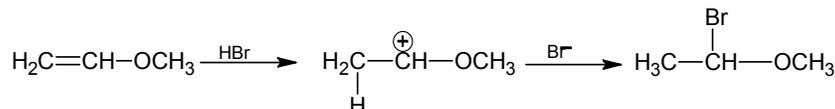


CHEMISTRY
PART – C

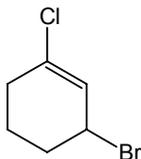
96. HBr reacts with $\text{CH}_2 = \text{CH} - \text{OCH}_3$ under anhydrous conditions at room temperature to give
 (1) CH_3CHO and CH_3Br (2) BrCH_2CHO and CH_3OH
 (3) $\text{BrCH}_2 - \text{CH}_2 - \text{OCH}_3$ (4) $\text{H}_3\text{C} - \text{CHBr} - \text{OCH}_3$

Ans. (4)

Sol. Electrophilic addition reaction more favourable.



97. The IUPAC name of the compound shown below is



- (1) 2-bromo-6-chlorocyclohex-1-ene (2) 6-bromo-2-chlorocyclohexene
 (3) 3-bromo-1-chlorocyclohexene (4) 1-bromo-3-chlorocyclohexene

Ans. (3)

98. The increasing order of the rate of HCN addition to compounds A – D is

- (A) HCHO (B) CH_3COCH_3
 (C) PhCOCH_3 (D) PhCOPh
 (1) $A < B < C < D$ (2) $D < B < C < A$
 (3) $D < C < B < A$ (4) $C < D < B < A$

Ans. (3)

99. How many moles of magnesium phosphate, $\text{Mg}_3(\text{PO}_4)_2$ will contain 0.25 mole of oxygen atoms?

- (1) 0.02 (2) 3.125×10^{-2}
 (3) 1.25×10^{-2} (4) 2.5×10^{-2}

Ans. (2)

Sol. $\text{Mg}_3(\text{PO}_4)_2$
 'n' moles

$$8n = 0.25$$

$$n = \frac{0.25}{8}$$

$$= \frac{25}{8 \times 100} = 3.125 \times 10^{-2}$$

100. According to Bohr's theory, the angular momentum of an electron in 5th orbit is

- (1) $25 \frac{h}{\pi}$ (2) $1.0 \frac{h}{\pi}$
 (3) $10 \frac{h}{\pi}$ (4) $2.5 \frac{h}{\pi}$

Ans. (4)

Sol. $mvr = \frac{nh}{2\pi}$
 $= \frac{5h}{2\pi} = 2.5 \frac{h}{\pi}$

101. Which of the following molecules/ions does not contain unpaired electrons?

- (1) O_2^{2-} (2) B_2
(3) N_2^+ (4) O_2

Ans. (1)

102. Total volume of atoms present in a face-centre cubic unit cell of a metal is (r is atomic radius)

- (1) $\frac{20}{3}\pi r^3$ (2) $\frac{24}{3}\pi r^3$
(3) $\frac{12}{3}\pi r^3$ (4) $\frac{16}{3}\pi r^3$

Ans. (4)

Sol. $V = n \times \left(\frac{4}{3}\pi r^3\right)$
 $= 4 \times \left(\frac{4}{3}\pi r^3\right)$
 $= \frac{16}{3}\pi r^3$

103. A reaction was found to be second order with respect to the concentration of carbon monoxide. If the concentration of carbon monoxide is doubled, with everything else kept the same, the rate of reaction will

- (1) remain unchanged (2) triple
(3) increase by a factor of 4 (4) double

Ans. (3)

Sol. $R \propto [W]^2$
 $R' \propto [2CO]^2$
 $R \propto 4[W]^2$
 $R \propto 4M$

104. Which of the following chemical reactions depicts the oxidizing behaviour of H_2SO_4 ?

- (1) $2HI + H_2SO_4 \longrightarrow I_2 + SO_2 + 2H_2O$ (2) $Ca(OH)_2 + H_2SO_4 \longrightarrow CaSO_4 + 2H_2O$
(3) $NaCl + H_2SO_4 \longrightarrow NaHSO_4 + HCl$ (4) $2PCl_5 + H_2SO_4 \longrightarrow 2POCl_3 + 2HCl + SO_2Cl_2$

Ans. (1)

105. The IUPAC name for the complex $[Co(NO_2)(NH_3)_5]Cl_2$ is

- (1) nitrito-N-pentaamminecobalt (III) chloride (2) nitrito-N-pentaamminecobalt (II) chloride
(3) pentaammine nitrito-N-cobalt (II) chloride (4) pentaammine nitrito-N-cobalt (III) chloride

Ans. (4)

106. The term anomers of glucose refers to

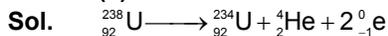
- (1) isomers of glucose that differ in configurations at carbons one and four (C-1 and C-4)
(2) a mixture of (D)-glucose and (L)-glucose
(3) enantiomers of glucose
(4) isomers of glucose that differ in configuration at carbon one (C-1)

Ans. (4)

FIITJEE Solutions to AIEEE–2006

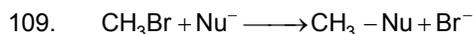
107. In the transformation of ${}_{92}^{238}\text{U}$ to ${}_{92}^{234}\text{U}$, if one emission is an α -particle, what should be the other emission(s)?
- (1) Two β^- (2) Two β^- and one β^+
 (3) One β^- and one γ (4) One β^+ and one β^-

Ans. (1)



108. Phenyl magnesium bromide reacts with methanol to give
- (1) a mixture of anisole and $\text{Mg}(\text{OH})\text{Br}$ (2) a mixture of benzene and $\text{Mg}(\text{OMe})\text{Br}$
 (3) a mixture of toluene and $\text{Mg}(\text{OH})\text{Br}$ (4) a mixture of phenol and $\text{Mg}(\text{Me})\text{Br}$

Ans. (2)



The decreasing order of the rate of the above reaction with nucleophiles (Nu^-) A to D is
 [$\text{Nu}^- = (\text{A}) \text{PhO}^-$, (B) AcO^- , (C) HO^- , (D) CH_3O^-]

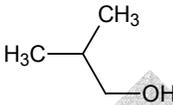
- (1) $\text{D} > \text{C} > \text{A} > \text{B}$ (2) $\text{D} > \text{C} > \text{B} > \text{A}$
 (3) $\text{A} > \text{B} > \text{C} > \text{D}$ (4) $\text{B} > \text{D} > \text{C} > \text{A}$

Ans. (1)

110. The pyrimidine bases present in DNA are
- (1) cytosine and adenine (2) cytosine and guanine
 (3) cytosine and thymine (4) cytosine and uracil

Ans. (3)

111. Among the following the one that gives positive iodoform test upon reaction with I_2 and NaOH is

- (1) $\text{CH}_3\text{CH}_2\text{CH}(\text{OH})\text{CH}_2\text{CH}_3$ (2) $\text{C}_6\text{H}_5\text{CH}_2\text{CH}_2\text{OH}$
 (3)  (4) PhCHOHCH_3

Ans. (4)

112. The increasing order of stability of the following free radicals is

- (1) $(\text{CH}_3)_2\dot{\text{C}}\text{H} < (\text{CH}_3)_3\dot{\text{C}} < (\text{C}_6\text{H}_5)_2\dot{\text{C}}\text{H} < (\text{C}_6\text{H}_5)_3\dot{\text{C}}$
 (2) $(\text{C}_6\text{H}_5)_3\dot{\text{C}} < (\text{C}_6\text{H}_5)_2\dot{\text{C}}\text{H} < (\text{CH}_3)_3\dot{\text{C}} < (\text{CH}_3)_2\dot{\text{C}}\text{H}$
 (3) $(\text{C}_6\text{H}_5)_2\dot{\text{C}}\text{H} < (\text{C}_6\text{H}_5)_3\dot{\text{C}} < (\text{CH}_3)_3\dot{\text{C}} < (\text{CH}_3)_2\dot{\text{C}}\text{H}$
 (4) $(\text{CH}_3)_2\dot{\text{C}}\text{H} < (\text{CH}_3)_3\dot{\text{C}} < (\text{C}_6\text{H}_5)_3\dot{\text{C}} < (\text{C}_6\text{H}_5)_2\dot{\text{C}}\text{H}$

Ans. (1)

113. Uncertainty in the position of an electron (mass = 9.1×10^{-31} kg) moving with a velocity 300 ms^{-1} , accurate upto 0.001%, will be

- (1) $19.2 \times 10^{-2} \text{ m}$ (2) $5.76 \times 10^{-2} \text{ m}$
 (3) $1.92 \times 10^{-2} \text{ m}$ (4) $3.84 \times 10^{-2} \text{ m}$
 ($h = 6.63 \times 10^{-34} \text{ Js}$)

Ans. (3)

Sol. $\Delta x \cdot \Delta V \geq \frac{h}{4\pi m}$

$$\Delta x \geq \frac{h}{4\pi m \Delta V} = \frac{6.63 \times 10^{-34}}{4 \times 3.14 \times 9.1 \times 10^{-31} \times 300 \times \frac{0.001}{100}}$$

$$= \frac{6.63 \times 10^{-34}}{4 \times 3.14 \times 9.1 \times 3 \times 10^{-31} \times 10^{-3}}$$

$$= 0.01933$$

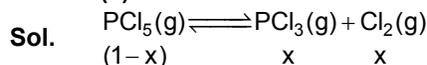
$$= 1.93 \times 10^{-2}$$

114. Phosphorus pentachloride dissociates as follows, in a closed reaction vessel,
 $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$

If total pressure at equilibrium of the reaction mixture is P and degree of dissociation of PCl_5 is x, the partial pressure of PCl_3 will be

- | | |
|-----------------------------------|------------------------------------|
| (1) $\left(\frac{x}{x+1}\right)P$ | (2) $\left(\frac{2x}{1-x}\right)P$ |
| (3) $\left(\frac{x}{x-1}\right)P$ | (4) $\left(\frac{x}{1-x}\right)P$ |

Ans. (1)



$$P_{\text{PCl}_3} = \left(\frac{x}{1+x}\right) \times P$$

115. The standard enthalpy of formation ($\Delta_f H^\circ$) at 298 K for methane, $\text{CH}_4(\text{g})$, is $-74.8 \text{ kJ mol}^{-1}$. The additional information required to determine the average energy for C – H bond formation would be

- (1) the dissociation energy of H_2 and enthalpy of sublimation of carbon
- (2) latent heat of vapourization of methane
- (3) the first four ionization energies of carbon and electron gain enthalpy of hydrogen
- (4) the dissociation energy of hydrogen molecule, H_2

Ans. (1)

116. Among the following mixtures, dipole-dipole as the major interaction, is present in

- | | |
|-------------------------|--------------------------------------|
| (1) benzene and ethanol | (2) acetonitrile and acetone |
| (3) KCl and water | (4) benzene and carbon tetrachloride |

Ans. (2)

117. Fluorobenzene ($\text{C}_6\text{H}_5\text{F}$) can be synthesized in the laboratory

- (1) by heating phenol with HF and KF
- (2) from aniline by diazotisation followed by heating the diazonium salt with HBF_4
- (3) by direct fluorination of benzene with F_2 gas
- (4) by reacting bromobenzene with NaF solution

Ans. (2)

118. A metal, M forms chlorides in its +2 and +4 oxidation states. Which of the following statements about these chlorides is correct?

- (1) MCl_2 is more volatile than MCl_4
- (2) MCl_2 is more soluble in anhydrous ethanol than MCl_4
- (3) MCl_2 is more ionic than MCl_4
- (4) MCl_2 is more easily hydrolysed than MCl_4

Ans. (3)

119. Which of the following statements is true?
 (1) H_3PO_3 is a stronger acid than H_2SO_3
 (2) In aqueous medium HF is a stronger acid than HCl
 (3) HClO_4 is a weaker acid than HClO_3
 (4) HNO_3 is a stronger acid than HNO_2

Ans. (4)

120. The molar conductivities $\Lambda_{\text{NaOAc}}^\circ$ and $\Lambda_{\text{HCl}}^\circ$ at infinite dilution in water at 25°C are 91.0 and $426.2 \text{ S cm}^2/\text{mol}$ respectively. To calculate $\Lambda_{\text{HOAc}}^\circ$, the additional value required is

- (1) $\Lambda_{\text{H}_2\text{O}}^\circ$ (2) $\Lambda_{\text{KCl}}^\circ$
 (3) $\Lambda_{\text{NaOH}}^\circ$ (4) $\Lambda_{\text{NaCl}}^\circ$

Ans. (4)

Sol. $\Lambda_{\text{CH}_3\text{COONa}}^\circ = \Lambda_{\text{CH}_3\text{COO}^-}^\circ + \Lambda_{\text{Na}^+}^\circ \dots\dots\dots (1)$

$\Lambda_{\text{HCl}}^\circ = \Lambda_{\text{H}^+}^\circ + \Lambda_{\text{Cl}^-}^\circ \dots\dots\dots (2)$

$\Lambda_{\text{NaCl}}^\circ = \Lambda_{\text{Na}^+}^\circ + \Lambda_{\text{Cl}^-}^\circ \dots\dots\dots (3)$

$\Lambda_{\text{CH}_3\text{COOH}}^\circ = (1) + (2) - (3)$

121. Which one of the following sets of ions represents a collection of isoelectronic species?

- (1) K^+ , Cl^- , Ca^{2+} , Sc^{3+} (2) Ba^{2+} , Sr^{2+} , K^+ , S^{2-}
 (3) N^{3-} , O^{2-} , F^- , S^{2-} (4) Li^+ , Na^+ , Mg^{2+} , Ca^{2+}

Ans. (1)

122. The correct order of increasing acid strength of the compounds

- (a) $\text{CH}_3\text{CO}_2\text{H}$ (b) $\text{MeOCH}_2\text{CO}_2\text{H}$
 (c) $\text{CF}_3\text{CO}_2\text{H}$ (d) $\begin{array}{c} \text{Me} \\ | \\ \text{C} - \text{CO}_2\text{H} \\ | \\ \text{Me} \end{array}$

is

- (1) $b < d < a < c$ (2) $d < a < c < b$
 (3) $d < a < b < c$ (4) $a < d < c < b$

Ans. (3)

123. In which of the following molecules/ions are all the bonds not equal?

- (1) SF_4 (2) SiF_4
 (3) XeF_4 (4) BF_4^-

Ans. (1)

124. What products are expected from the disproportionation reaction of hypochlorous acid?

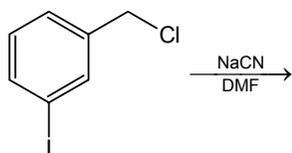
- (1) HClO_3 and Cl_2O (2) HClO_2 and HClO_4
 (3) HCl and Cl_2O (4) HCl and HClO_3

Ans. (4)

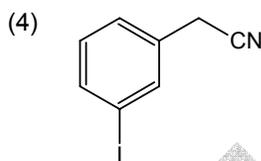
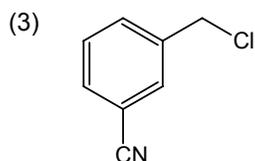
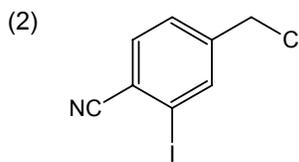
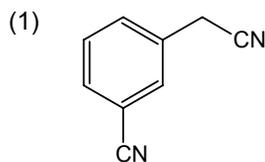
125. Nickel ($Z = 28$) combines with a uninegative monodentate ligand X^- to form a paramagnetic complex $[\text{NiX}_4]^{2-}$. The number of unpaired electron(s) in the nickel and geometry of this complex ion are, respectively

- (1) one, tetrahedral (2) two, tetrahedral
 (3) one, square planar (4) two, square planar

131. The structure of the major product formed in the following reaction



is



Ans. (4)

132. Reaction of trans-2-phenyl-1-bromocyclopentane on reaction with alcoholic KOH produces

- (1) 4-phenylcyclopentene (2) 2-phenylcyclopentene
 (3) 1-phenylcyclopentene (4) 3-phenylcyclopentene

Ans. (4)

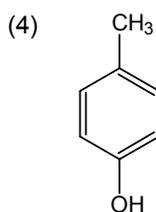
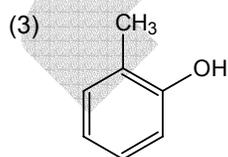
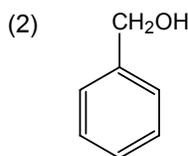
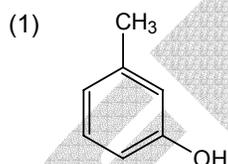
Sol. According to E₂ mechanism.

133. Increasing order of stability among the three main conformations (i.e. Eclipse, Anti, Gauche) of 2-fluoroethanol is

- (1) Eclipse, Gauche, Anti (2) Gauche, Eclipse, Anti
 (3) Eclipse, Anti, Gauche (4) Anti, Gauche, Eclipse

Ans. (3)

134. The structure of the compound that gives a tribromo derivative on treatment with bromine water is

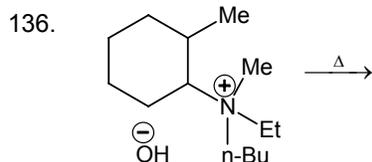


Ans. (1)

135. The decreasing values of bond angles from NH₃ (106°) to SbH₃ (101°) down group-15 of the periodic table is due to

- (1) increasing bp-bp repulsion (2) increasing p-orbital character in sp³
 (3) decreasing lp-bp repulsion (4) decreasing electronegativity

Ans. (4)



The alkene formed as a major product in the above elimination reaction is

- (1)
- (2) $\text{CH}_2 = \text{CH}_2$
- (3)
- (4)

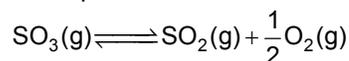
Ans. (2)

137. The “spin-only” magnetic moment [in units of Bohr magneton, (μ_B)] of Ni^{2+} in aqueous solution would be (Atomic number of Ni = 28)

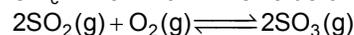
- (1) 2.84 (2) 4.90
 (3) 0 (4) 1.73

Ans. (1)

138. The equilibrium constant for the reaction



is $K_c = 4.9 \times 10^{-2}$. The value of K_c for the reaction



will be

- (1) 416 (2) 2.40×10^{-3}
 (3) 9.8×10^{-2} (4) 4.9×10^{-2}

Ans. (1)

Sol.
$$K'_c = \left(\frac{1}{4.9 \times 10^{-2}} \right)^2$$

$$= \frac{10^4}{4.9 \times 4.9} = \frac{100 \times 100}{24.01}$$

$$= 4.1649 \times 100$$

$$= 416.49$$

139. Following statements regarding the periodic trends of chemical reactivity of the alkali metals and the halogens are given. Which of these statements gives the correct picture?

- (1) The reactivity decreases in the alkali metals but increases in the halogens with increase in atomic number down the group
- (2) In both the alkali metals and the halogens the chemical reactivity decreases with increase in atomic number down the group
- (3) Chemical reactivity increases with increase in atomic number down the group in both the alkali metals and halogens
- (4) In alkali metals the reactivity increases but in the halogens it decreases with increase in atomic number down the group

Ans. (4)

Ans. (4)

Sol. There is one mistake in Question paper.

Assuming concentration of solution is 0.2 M instead of 0.02 M. Since resistance of 0.2 M is 520 Ω.

$$R = 100 \Omega$$

$$K = \frac{1}{R} \left(\frac{\ell}{a} \right)$$

$$1.29 = \frac{1}{100} \left(\frac{\ell}{a} \right)$$

$$\left(\frac{\ell}{a} \right) = 129 \text{ m}^{-1}$$

$$R = 520 \Omega, C = 0.2 \text{ M}$$

$$K = \frac{1}{R} \left(\frac{\ell}{a} \right) = \frac{1}{520} (129) \Omega^{-1} \text{m}^{-1}$$

$$\mu = K \times V_{\text{in cm}^3}$$

$$= \frac{1}{520} \times 129 \times \frac{1000}{0.2} \times 10^{-6} \text{ m}^3$$

$$= \frac{129}{520} \times \frac{1000}{0.2} \times 10^{-6}$$

$$= 1.24 \times 10^{-3}$$

$$= 12.4 \times 10^{-4}$$

144. The ionic mobility of alkali metal ions in aqueous solution is maximum for

- | | |
|---------------------|---------------------|
| (1) K ⁺ | (2) Rb ⁺ |
| (3) Li ⁺ | (4) Na ⁺ |

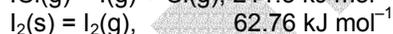
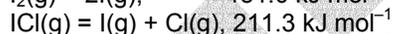
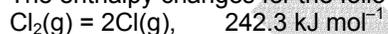
Ans. (2)

145. Density of a 2.05 M solution of acetic acid in water is 1.02 g/mL. The molality of the solution is

- | | |
|-------------------------------|-------------------------------|
| (1) 1.14 mol kg ⁻¹ | (2) 3.28 mol kg ⁻¹ |
| (3) 2.28 mol kg ⁻¹ | (4) 0.44 mol kg ⁻¹ |

Ans. (3)

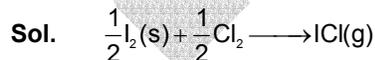
146. The enthalpy changes for the following processes are listed below:



Given that the standard states for iodine and chlorine are I₂(s) and Cl₂(g), the standard enthalpy of formation for ICl(g) is

- | | |
|--------------------------------|---------------------------------|
| (1) -14.6 kJ mol ⁻¹ | (2) -16.8 kJ mol ⁻¹ |
| (3) +16.8 kJ mol ⁻¹ | (4) +244.8 kJ mol ⁻¹ |

Ans. (3)



$$\Delta H = \left[\frac{1}{2} \Delta H_{\text{I}_2(\text{s}) \rightarrow \text{I}_2(\text{g})} + \frac{1}{2} \mu_{\text{I-I}} + \frac{1}{2} \mu_{\text{Cl-Cl}} \right] - [\mu_{\text{I-Cl}}]$$

$$= \left(\frac{1}{2} \times 62.76 + \frac{1}{2} \times 151.0 + \frac{1}{2} \times 242.3 \right) - (211.3)$$

$$= 228.03 - 211.3$$

$$\Delta H = 16.73$$

147. How many EDTA (ethylenediaminetetraacetic acid) molecules are required to make an octahedral complex with a Ca²⁺ ion?

- | | |
|---------|-----------|
| (1) Six | (2) Three |
| (3) One | (4) Two |

Ans. (3)

